Structure and bonding of [(SIPr)AgX] (X = Cl, Br, I and OTf)†

Valerie H. L. Wong, Andrew J. P. White, T. S. Andy Hor and King Kuok (Mimi) Hii*

A series of iso-structural complexes [(SIPr)AgX] (X = Cl, Br, I, OTf; SIPr = 1,3-bis(2,6-diisopropylphenyl)imidazolidene) were synthesised, including the first example of a N-heterocyclic carbene silver(I) complex containing an O-bound triflate. Bond Energy Dissociation and Natural Orbitals for Chemical Valence bond analyses (BEDA & ETS-NOCV) revealing a significant NHC→M σ-back-donation, which influences the stability and sigma-donicity of these complexes.

The discovery of a stable free N-heterocyclic carbene (NHC) by Arduengo in 1991† is an important milestone that has transformed many areas of chemistry. In a little over two decades, NHC’s have been shown to be a privileged ligand class for important metal-catalysed reactions with broad applicability, such as cross-coupling and olefin metathesis, as well as being active organocatalysts in their own right. Additionally, metal–NHC complexes have also found applications as valuable functional materials, and anti-microbial and anti-tumour agents for the treatment of multi-drug resistant infection and cancer. Interests in their chemical, physical and biological activities have stimulated many studies of the metal–carbene bond.

Monomeric Ag(i)–NHC complexes with the structure [NHC]AgX, where X = anion ligand, are relatively uncommon. Generally, a combination of strong σ-donor ability of the NHC ligand and a relatively weak Ag–X bond makes these complexes unstable; they tend to undergo rapid exchange in solution with the homoleptic species [NHC]₂Ag[AgX]₂, which may be isolated as ionic pairs, or as discreet bis-NHC complexes [NHC]₂Ag[X] (when X = counter-anion, e.g. BF₄⁻). Concurrently, [NHC]AgX complexes where NHC = imidazolidin-2-ylidene derivatives (‘saturated carbenes’) are even less common. Saturated NHC ligands are generally thought to be better σ-donors than their unsaturated analogues, thus can impose greater trans-effect on the opposite anion, making these complexes even more unstable.

To date, there are no reported examples of mononuclear [NHC]AgX complexes containing a bound triflate anion (X = CF₃SO₃⁻), although the synthesis of [(SIPr)AgOTf] (1) had been reported to be prepared by the addition of AgOTf to a suspension of [(SIPr)AgCl]. Within our work programme on silver catalysis, the procedure was repeated to afford a white crystalline solid that has similar characterisation data to that described: the key observation being a downfield shift of the 13C NMR resonance signal corresponding to the metal-coordinated carbene to 205.3 ppm, with J coupling to 109Ag/107Ag isotopes of 304 and 264 Hz, respectively. Subsequently, a crystal suitable for X-ray crystallography was obtained, revealing that it is, in fact, a chloride-bridged dinuclear complex, [(SIPr)AgCl₂]OTf (2) (Fig. 1), consisting of two [Ag(SIPr)]⁺ units connected by a bridging chloride, structurally analogous to a fluoride-bridged dimer [(SIPr)Ag₂(μ-F)][BF₄⁻] reported recently. The Ag–C bond distance of 2.08 Å in complex 2 is longer than that found in the fluoride congener (2.05 Å), reflecting the greater trans-effect of Cl. Interestingly, very similar 13C NMR data were reported for the two complexes, suggesting that the halide bridges in these complexes may not be retained in solution.

Fig. 1 X-ray crystal structure of compound 2 (H atoms omitted for clarity).

† Electronic supplementary information (ESI) available: Synthesis procedure, characterisation data, including X-ray data, computational details and NMR spectra. CCDC 1400409–1400412. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5cc07977b
Addition of further equivalents of AgOTf to complex 2 did not lead to further halide abstraction. This led us to explore the use of the corresponding bromide and iodide precursors (3 and 4, respectively) for the metathesis reaction. These were prepared separately from different imidazolinium salts and fully characterized as monomeric [(NHC)AgX] complexes (Scheme 1), as confirmed by X-ray crystallography (see ESIF).

The addition of AgOTf to either 3 or 4 led to the formation of a new compound that is distinct from that obtained from the chloride precursor. In this case, the product was found to be highly fluxional in solution. The structure of [(SIPr)AgOTf] (1) was eventually confirmed by X-ray diffraction study (Fig. 2), showing that Ag is coordinated in a linear arrangement to SIPr and the triflate anion, providing the first unequivocal evidence that a Ag–O bond can be forged between a triflate anion and (NHC)–silver(i). The stability of 1 is both unique and surprising, as the corresponding experiments with IPr (unsaturated carbene) and SIMes analogues (where i-Pr is substituted with Me) resulted in the formation of homoleptic [(NHC)2Ag] complexes (Scheme 1).

While the latter can be attributed to greater kinetic lability due to the smaller ligand, the former suggests that electronic factors are also important.

In order to understand this, the electronic structures of these iso-structural series of complexes were investigated. Bond energy decomposition analysis (BEDA)\textsuperscript{13} is a popular method used for the analysis of NHC–metal complexes by partitioning the M–C bond into ionic and covalent components.\textsuperscript{46,14} This was employed by Frenking and co-workers in a computational study of [(NHC)AgX] complexes, where the NHC = unsubstituted imidazol-2-ylidine.\textsuperscript{15}

The results of our study are summarized in Table 1. In agreement with previous studies, the metal–carbene bond is largely ionic in nature, dominated by repulsive $\Delta E_{\text{Pauli}}$ and attractive $\Delta E_{\text{elstat}}$ terms, with a small, but not insignificant, orbital contribution. The binding energies between SIPr and AgX decrease slightly in the order: Cl > Br > I; this is supported experimentally by a corresponding increase in the Ag–C bond length, downfield shift of the $^{13}$C resonance signal, and reduced $J^{(\text{Ag–C})}$ coupling constants. In comparison, the Ag–C bond in the triflate complex is notably stronger than the halide compounds, due to an increase in both the ionic (steric int) and covalent ($\Delta d_{\text{covalent}}$) components. This is also supported by the shortest Ag–C bond.

| Table 1 | Bond energy decomposition analysis of Ag–C bond in [(SIPr)AgX] complexes (kcal mol$^{-1}$), experimental and calculated Ag–C bond distances (Å), $\delta C$ (ppm) and $J^{(\text{Ag–C})}$ coupling constants (Hz) |
| --- | --- | --- | --- | --- | --- |
| Energies (kcal mol$^{-1}$) | X = Cl & 39.21 & 74.36 & 61.46 & 58.20 |
| $\Delta E_{\text{int}}$ | $\Delta E_{\text{int}} + \Delta E_{\text{electrost}} + \Delta E_{\text{Pauli}}$ | -50.71 & -49.28 & -47.37 & -59.20 |
| Steric int | -13.68 | -12.22 | -11.11 & -16.30 |
| $\Delta E_{\text{elstat}}$ | 120.04 & 119.96 & 115.60 & 126.71 |
| $\Delta E_{\text{Pauli}}$ | -133.72 & -132.19 & -126.71 & -143.01 |
| $\Delta E_{\text{Pauli}}$ | -37.04 | -37.06 & -36.26 & -42.90 |
| Ag–C (expt) | 2.081(9) & 2.080(3) & 2.094(3) & 2.072(3) |
| Ag–C (calc) | 2.071 & 2.083 & 2.099 & 2.058 |
| $\delta C$ (ppm) | 207.7 | 209.9 & 210.8 & — |
| $J^{(\text{Ag–C})}$ | 253, 219 & 249, 216 & 238, 207 & — |

a $\Delta E_{\text{int}} = \Delta E_{\text{int}} + \Delta E_{\text{electrost}} + \Delta E_{\text{Pauli}}$; sterical int. = $\Delta E_{\text{Pauli}} + \Delta E_{\text{electrost}}$.

b Resonance signal of coordinated carbene. Recorded in CD$_2$Cl$_2$.

c Values corresponding to $J^{(109\text{Ag–13C})}$ and $J^{(107\text{Ag–13C})}$ respectively.

d Not observed (fluxional).

Fig. 2 X-ray crystal structure of [(SIPr)AgOTf], 1. H atoms omitted for clarity. SIPr = 1,3-bis(2,6-disopropylphenyl)imidazolidene.
X = OTf ($\text{X} = \text{I}$), $\text{C} = \text{Br}$, ruthenium(II), palladium(II) and platinum(II) complexes, corresponding carbenes. In NMR and calorimetric studies of salts is not an adequate indicator for the strength of the CO ligands. In this study, however, we have shown that the σ-donor ability of saturated and unsaturated NHC ligands has been a subject of some debate. It is generally accepted that the pKa of the imidazolium/imidazolinium salts is not an adequate indicator for the σ-donory of their corresponding carbenes. In NMR and calorimetric studies of ruthenium(ii), palladium(ii) and platinum(ii) complexes, unsaturated carbenes (IPr and IMes) were reported to be less σ-donating than their saturated analogues (SIPr and SIMes). However, the reverse appears to be true in IR spectroscopic studies of nickel(0) and iridium(iii) carbonyl complexes; explained by evoking synergistic relationships between NHC and CO ligands. In this study, however, we have shown that SIPr can be a weaker σ-donor than IPr in the absence of other σ-acceptor ligands. For Ag(i) complexes, the presence of a full-shell of electrons (d10) prevents effective π- and π- back donation, particularly as saturated carbenes are poor π-acceptors. Consequently, the excess electron density on the metal is dispelled back towards the metal via σbd, reducing the σ-donicity of the NHC ligand, and correspondingly, its trans-effect on the opposite anion.

This electronic ‘levelling’ effect is bound to have important implications in the application of these metal complexes as catalysts. Given that the [(IPr)Ag(Otf)] cannot be prepared, preliminary catalytic experiments were performed using SIPr and IPr complexes of silver(i) benzoate so as to afford direct comparisons. Nearly identical rates in the cyclization of propargyl amide 5 to the oxazolidine 6 were observed (Fig. 4), supporting the view that the two complexes are not electronically different.

In summary, a series of monomeric [(SIPr)AgX] compounds have been prepared and fully characterised; including the first report of a triflate-coordinated complex. BDEA showed that the Ag–C binding energy increases in the order of X = I < Br < Cl < OTf, while ETS-NOCV analyses revealed significant involvement of a σ-component in the back donation of electron density from Ag → NHC, making the SIPr ligand a weaker σ-donor than its saturated analogues in these complexes. This implies that, at least for group 11 metals, that the choice between saturated or unsaturated NHC ligands may not make much of a difference in their catalytic activities as may be expected.

VHLW was supported by a NUS-Imperial Scholarship. We are grateful to Dr Amanda Simperler and Dr Helen Tsui (EPSRC UK National Service for Computational Chemistry Software, NSCCS) for help with ADF calculations.

Notes and references
ChemComm

Communication

4 (a) H. Jacobsen, A. Correa, A. Poater, C. Costabile and L. Cavallo,
Coord. Chem. Rev., 2009, 253, 687; (b) J. C. Bernhammer, G. Frison
and H. V. Huynh, Chem. – Eur. J., 2013, 19, 12892; (c) D. J. Nelson
6 H.-L. Su, L. M. Pérez, S.-J. Lee, J. H. Reinenspies, H. S. Bazzi and
8 F. Lazreg, D. B. Cordes, A. M. Z. Slawin, C. S. J. Cazin, Organo-
metallics, 2015, 34, 419.
9 CCDC search revealed only 10 examples: X = Cl (7 e.g.);
(a) O. Winkelmann, C. Nüther and U. Liming, J. Organomet. Chem.,
2008, 693, 923; (b) P. de Frémont, N. M. Scott, E. D. Stevens,
T. Ramnial, O. C. Lightbody, C. L. Cannon and W. J. Youngs,
(f) T. Makino, R. Yamasaki, I. Azamaya, H. Masu and S. Saito, Organo-
metallics, 2010, 29, 6291. X = Br (1 e.g.): (g) J. Pykowicz, S. Roland and
10 B. K. Tate, C. M. Wyss, J. Bacsa, M. Wieliczko and J. P. Sadighi,
Chem. Commun., 2013, 49, 9272; (b) J. L. Arbour, H. S. Rzepa, J. Contreras-Garcia, L. A. Adrio,
11 (a) V. H. L. Wong, T. S. A. Hor and K. K. Hii, Chem. Commun., 2013,
49, 9272; (b) J. L. Arbour, H. S. Rzepa, J. Contreras-Garcia, L. A. Adrio,
12 C. M. Wyss, B. K. Tate, J. Bacsa, M. Wieliczko and J. P. Sadighi,
Polyhedron, 2014, 84, 87.
15 D. Nemešok, K. Wichmann and G. Frenking, Organometallics, 2004,
23, 3640.
16 (a) D. Marchione, L. Belpassi, G. Bistoni, A. Macchioni, F. Tarantelli
and D. Zuccaccia, Organometallics, 2014, 33, 4200; (b) L. Batiste and
17 Presumably because SIPr ligand (a weaker σ-donor) exerts a smaller
trans-influence than IPr.
18 The reaction has been shown to be sensitive to ligand electronic
effects, see: (a) V. H. L. Wong, A. J. P. White, T. S. A. Hor and
(b) V. H. L. Wong and K. K. Hii, unpublished results.
19 A. C. Hillier, W. J. Sommer, B. S. Yong, J. L. Petersen, L. Cavallo and
20 H. V. Huynh, Y. Han, R. Jothibasu and J. A. Yang, Organometallics,
2009, 28, 5395.
21 S. Fantasia, J. L. Petersen, H. Jacobsen, L. Cavallo and S. P. Nolan,
Organometallics, 2007, 26, 5880.
22 R. Dorta, E. D. Stevens, N. M. Scott, C. Costabile, L. Cavallo,
23 R. A. Kelly, H. Clavier, S. Giudice, N. M. Scott, E. D. Stevens,
J. Bordner, I. Samardjiev, C. D. Hoff and S. P. Nolan,
24 Overall catalytic turnover is determined by a push–pull effect,
constituting the σ-donicity of the ligand to labilise the anion ligand,
and the π-acidity of the metal centre, in this case, to activate the
unsaturated bond towards nucleophilic attack.

View Article Online

This journal is © The Royal Society of Chemistry 2015

Open Access Article. Published on 16 October 2015. Downloaded on 23/10/2015 11:56:27.
This article is licensed under a
Creative Commons Attribution 3.0 Unported Licence.