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Azoniaspiro salts: towards bridging the gap between room-temperature ionic liquids and molten salts†

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In a continued effort to improve the suitability of ionic liquids in applications operating at raised temperatures, novel spirocyclic 'azoniaspiro' salts (with cations derived from five-, six-, seven- and eight-membered rings) are prepared and characterised. The structural and thermal properties of these salts are compared against those of established analogues. The stable geometries and ion pairing behaviour of these species are investigated *via* a combined experimental/computational approach, employing X-ray crystallography and Density Functional Theory (DFT) methods. Subsequently, the thermal stabilities of these organic salts are characterised and compared using a broad range of techniques. Hyphenated Thermogravimetry-Mass Spectrometry investigations enable complex mechanisms underlying thermal decomposition to be elucidated. Lastly, transition state structures are optimised, corresponding to plausible decomposition mechanisms of the azoniaspiro salt, 6-azoniaspiro[6.5]dodecanium chloride, and one prototypical monocyclic species 1-butyl-1-methylpiperidinium chloride, using DFT. The observed improved thermal stabilities of the azoniaspiro salts, and their potential higher-temperature stable-liquid ranges, render them promising candidates for future ionic liquid applications.

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Introduction

Ionic liquids are species constituted entirely of ions, that exist in the molten state at or marginally above room temperature.¹ Primitive ionic liquids belonged to a restricted class of air-/moisture-sensitive *N*-alkylpyridinium halometallate 'melts'.^{2–7} By contrast, contemporary room-temperature ionic liquids are stable in the presence of both moisture and oxygen; they are commonly derived from alternative unsaturated heterocycles (*e.g.* imidazole,⁸ 1,8-diazabicyclo[5.4.0]undec-7-ene⁹) or saturated heterocycles (*e.g.* pyrrolidine,¹⁰ piperidine), and incorporate a molecular anion with well-defined structure. Dialkylpiperidinium ionic liquids have been identified as strong candidate electrolytes for battery technology, attributable to their wide electrochemical windows.¹¹ The seven-membered heterocycle azepane is a by-product of polyamide synthesis; it has been suggested, therefore, that dialkylazepanium ionic liquids may represent more sustainable analogues of pyrrolidine- and piperidine-derived ionic liquids.¹²

One common feature of the above-described cation classes is that, in each case, the positive charge is associated with a quaternary nitrogen centre. At present under-investigated, 'azoniaspiro'-type salts represent an interesting additional potential class of ionic liquid cation, whereby two rings are fused together *via* a quaternary nitrogen atom.¹³ The presence of the two cyclic moieties offers a greater variety of possible structures, arising from combinations of normal to medium sized rings, and the substitution of heteroatoms into either or both of the rings. In the late 1970s, azoniaspiro salts were investigated as 'soluble hydrophobic species'.¹⁴ Tetraalkyl-ammonium salts incorporating monocyclic and azoniaspiro cations have played the role of 'structure-directing agents' in zeolite synthesis.^{15–17} More recently, mixtures of azoniaspiro salts and ethylene carbonate (EC) have been investigated as electrolytes for use in Li batteries.¹⁸ To the authors knowledge, azoniaspiro salts have not yet been widely explored in the context of ionic liquid chemistry.

For any application involving a tetraalkylammonium salt at elevated temperatures, an understanding of the thermal stability limitations of this specific salt is crucial. This must include knowledge of the safe operating temperatures, the underlying chemical mechanisms of degradation, and the properties of all thermal decomposition products formed. The measurement of these thermal properties requires a combination of experimental and computational techniques.

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Thermogravimetric analysis (TGA) has been widely implemented in the evaluation of ionic liquid thermal stabilities.^{19–30} Temperature-ramped TGA experiments measure the changes in weight% as a function of temperature, whilst maintaining a constant heating rate. Heating rates of 10 °C min⁻¹ or 20 °C min⁻¹ are typical. However, caution must be exercised to keep experimental conditions as uniform as possible when comparing any individual thermal stability parameter; the ' T_{onset} ' parameter allows rapid comparison of compounds in a series, where care has been taken to keep experimental conditions the same.²⁰ However, while it has been commonly used to assess ionic liquid stability, T_{onset} has been shown to greatly overestimate stability.^{21,22} Alternatively, ' T_{start} ' quantifies the first point of appreciable weight loss ($dw/dt \neq 0$). Isothermal TGA experiments provide insights into the long-term thermal stability, and Arrhenius parameters of thermal degradation.³¹

Herein we report syntheses and thermal characterisations of a broad series of tetraalkylammonium chloride salts, **1–10**, incorporating either one or two saturated rings of five, six, seven or eight atoms (Fig. 1). Thermal decomposition of ionic liquids frequently involves nucleophilic attack of the anion on pendant alkyl chains of the cation.^{28,32–35} It was therefore proposed that the inclusion of a second cyclic structure would enhance steric interactions around the nitrogen centre and α -carbons, conferring improved thermal stability on the salt.

For clarity, azoniapiro chloride salts included in this study are assigned abbreviated names according to the total number of ring atoms, including the quaternary nitrogen, *e.g.* 5-azoni-aspiro-[4.4]nonanium chloride (**4**) is '[5+5]Cl'. Compound **6**, incorporating the oxygen-substituted ring, is denoted '[3-oxa-6+6]Cl'. Monocyclic salts, **1–3**, are instead named according to established literature convention, for example '[C_nC_m]pyrr]Cl', where integers ' n '/' m ' denote the number of carbon atoms in the alkyl chains.¹

The structural properties of selected salts were investigated using X-ray crystallography and Density Functional Theory (DFT). Relative thermal stabilities and decomposition thermodynamics were compared by TGA. Furthermore, the thermal decomposition mechanisms of [C_4C_4]pip]Cl, **2**, and [7+6]Cl, **8** were studied using hyphenated thermogravimetric techniques (Thermogravimetric Analysis-Mass Spectrometry 'TGA-MS', and Thermogravimetric Analysis-Gas Chromatography/Mass Spectrometry 'TGA-GCMS'); degradation mechanisms were subsequently rationalised using DFT.

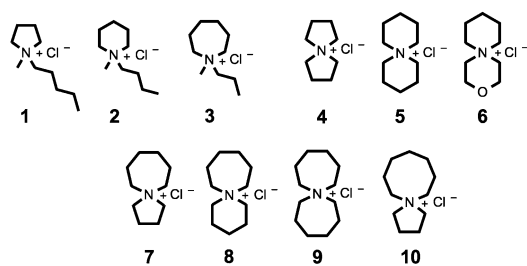


Fig. 1 Investigated single-ring (**1–3**) and azoniapiro (**4–10**) tetraalkylammonium chloride salts incorporating five, six, seven and eight-membered saturated rings.

'Room-Temperature Ionic Liquids' (RTILs) have been of particular interest, being in the molten state at, or below room temperature. However, RTILs often suffer from poor thermal stability, limiting utility at higher temperatures.^{36,37} Fully-inorganic molten salts are stable at elevated temperatures, but have high melting points that limit their utility at lower temperatures. It was proposed that azoniapiro salts may partially bridge this gap with respect to the stable-liquid range.

One common degradation pathway for ionic liquids is the nucleophilic attack of the anion on the cation.^{28,32–35} Since we were interested in the contributions of an array of azoniapiro cations to the relative stabilities of ionic liquids, we selected the chloride anion due to its ability to act as a nucleophile, the relative ease of crystallisation (enabling crystal structure data to be obtained) and for simplicity in conducting a computational survey of low-energy conformers. Nevertheless, for one azoniapiro salt, metathetic reactions were performed so as to exchange chloride for bulky, charge-diffuse, non-nucleophilic anions (triflate and bis(trifluoromethanesulfonyl)imide), with the intention of drastically lowering the melting points and therefore demonstrating the potential of using simple azonia-spiro salts, or mixtures, at (close to) ambient temperatures.

Experimental

The salts included in this study are shown in Fig. 1. All compounds were prepared from commercial starting materials. Single-ring chloride salts, **1–3**, were prepared by alkylation of an *N*-methylheterocycle with the relevant 1-chloroalkane according to established literature protocols.¹ *N*-methylazepane was prepared from commercial hexamethyleneimine (azepane) *via* Escheweiler–Clarke methylation with formic acid and formaldehyde.³⁸ Azoniapiro salts, **4–10**, were prepared by heating under reflux of a biphasic toluene–water solution, containing a nitrogen heterocycle (*i.e.* piperidine, morpholine, azepane or heptamethyleneimine), a 1,*n*-dichloroalkane, and one molar equivalent of sodium hydroxide. The full synthesis of [7+6]Cl, **8**, is described below. Synthesis and characterisation of **1–10** and intermediate compounds is described fully in the ESI.†

Thermogravimetric Analysis (TGA) experiments were performed with PerkinElmer 'Pyris 1' TGA apparatus. Temperature-ramped TGA experiments were carried out using 4–8 mg of the salt, with a heating rate of 10 °C min⁻¹, and employing platinum pans. In all circumstances, nitrogen was used as the inert carrier gas at a flow rate of 20 ml min⁻¹. Isothermal TGA was performed using 7–8 mg of the compound. Thermogravimetric Analysis-Mass Spectrometry (TGA-MS) experiments were performed on 20–60 mg of the salt, employing ceramic pans. A heating rate of 10 °C min⁻¹ and a CP Grade helium flow of 20 ml min⁻¹ were employed. 'Pyris 1' TGA apparatus was connected to a 'Hiden Analytical HPR 20' mass spectrometer *via* a ceramic heated capillary. Electron Ionisation (EI) mass spectrometry was used, with 70 eV ionisation energy. TGA, TGA-MS and TGA-GCMS experimental conditions are fully described in the ESI.† All calculations were performed at the B3LYP-D2/6-311++G(d,p)



level of theory. ΔG , ΔH , $T\Delta S$ and E (zero-point-corrected) energies are provided in the ESI†

Synthesis of 6-azoniaspiro[6.5]dodecanium chloride (8)

Azepane (15.0 g, 151 mmol) was diluted with toluene (70 ml) in a 250 ml, three-necked round-bottomed flask, fitted with a reflux condenser, a dropping funnel and a thermometer. Separately, an aqueous solution of NaOH was prepared by dissolving solid pellets of NaOH (6.05 g, 151 mmol) in deionized water (20 ml). The NaOH solution was carefully added to the toluene solution, and the dropping funnel was charged with 1,5-dichloropentane (21.33 g, 151 mmol). The biphasic mixture was heated to 80 °C under a nitrogen atmosphere and the 1,5-dichloropentane was added dropwise to the toluene/water mixture over a period of 15 minutes, with vigorous stirring. The solution was maintained at 80 °C for four days and was subsequently allowed to cool to room temperature. Toluene (100 ml) and deionized water (100 ml) were added, and the two layers were separated. The (upper) toluene layer was washed with deionized water (2 × 25 ml), and the three aqueous fractions were combined. Water was removed by rotary evaporation. To the resultant crude residue was added cold ethanol (200 ml), in which the product was found to be more soluble than NaCl. The solution was filtered, and the filtrate was treated with anhydrous MgSO₄. The solution was filtered again, and the ethanol was removed by rotary evaporation, to yield 6-azoniaspiro[6.5]dodecanium chloride (18.6 g, 60%) as a white, free-flowing solid. Found: m.p. 272–273 °C with decomposition (lit. 270–272 °C with decomposition).¹³ ¹H NMR (400 MHz, DMSO-d₆): δ 3.44 (4H, t, J = 5 Hz), 3.34 (4H, t, J = 6 Hz), 1.88–1.68 (8H, m), 1.65–1.47 (6H, m). ¹³C {¹H} NMR (100 MHz, DMSO-d₆): δ 61.4, 60.3, 27.1, 20.8, 20.3, 19.4. ν (neat)/cm⁻¹ 2933, 2862 (aliphatic C–H stretch, w), 1449 (aliphatic CH₂ bend, w), 870 (aliphatic C–N stretch, m). m/z (LSIMS⁺): 168 (100%) [(CH₂)₆N(CH₂)₅]⁺. m/z (LSIMS⁻): 35 (100%) ³⁵Cl⁻, 37 (23%) ³⁷Cl⁻. Calc. for C₁₁H₂₂ClN: C, 64.84; H, 10.88; N, 6.87%. Found: C, 64.81; H, 10.77; N, 6.89%.

Results and discussion

Initially, for selected compounds within the series (Fig. 1), a combination of X-ray crystallography and DFT techniques was employed in order to evaluate the preferential sites of the chloride anion around the tetraalkylammonium cation. Knowledge of the low-energy chloride bonding sites, and the stable geometries of single-ring and azoniaspiro cations, is valuable when proposing thermal decomposition mechanisms (and, in addition, when using these salts as structure-directing agents in zeolite synthesis). Moreover, structural information of this type allows for the prediction of possible solvating properties of the azoniaspiro salts.

The gas phase ion pairing behaviour of dialkylimidazolium chloride ionic liquids has been extensively investigated.^{39–42} Hunt and co-workers demonstrated that for the prototypical ionic liquid 1-butyl-3-methylimidazolium chloride, [C₄C₁im]Cl, Cl⁻ preferentially interacts with both the acidic C² proton on

the dialkylimidazolium cation and with an α -proton on either the methyl or butyl side chains.³⁹ More recently ‘dimer’ structures (stable clusters incorporating two cations and two anions) of [C₁C₁im]Cl have been investigated in detail, yielding a complex array of geometries.^{43–45}

The fully-saturated tetraalkylammonium chloride salts included in this investigation lack the acidic aromatic C² proton of the dialkylimidazolium analogues; the low-energy ion pair geometries are therefore anticipated to differ substantially. Single, colourless crystals of [5+5]Cl, 4, [6+6]Cl, 5, [7+5]Cl, 7, and [7+6]Cl, 8, were grown *via* the vapour diffusion of *n*-pentane (in which the salts were largely insoluble) into a saturated isopropanol solution of the azoniaspiro salt. For salt 5, containing two six-membered rings, an X-ray structure was obtained (shown in Fig. 2, Fig. E6 and E7, ESI†).

The X-ray crystal structure of 5 (Fig. 2) indicates that both six membered rings adopt a ‘chair’-type conformation around the central nitrogen atom of the azoniaspiro cation. Chloride anions are then preferentially located in multiple ‘tripod’-like arrangements interacting with three α (or β) protons on nearby azoniaspiro cations. The unit cell structure of 5 highlights an alternating array of cations and anions, each chloride anion interacting with tripods of α and β hydrogen atoms on each of four neighbouring cations (C–H...Cl distances 2.75–2.95 Å).

In the large crystals of [5+5]Cl, 4, the azoniaspiro salt was found to have co-crystallised with H₂O. A crystal structure of 4·H₂O is displayed in the ESI† (Fig. E5). Moreover, the crystalline material was found to be highly hygroscopic, and deliquesced rapidly when exposed to air. Attempts to remove the water by drying of an ethanol solution of 4 with MgSO₄ and subsequent recrystallisation of the sample proved unsuccessful. Therefore compound 4 is omitted from further discussion.

For salts 7 and 8 (containing a seven-membered ring), X-ray crystal structures were difficult to obtain. The scattering patterns

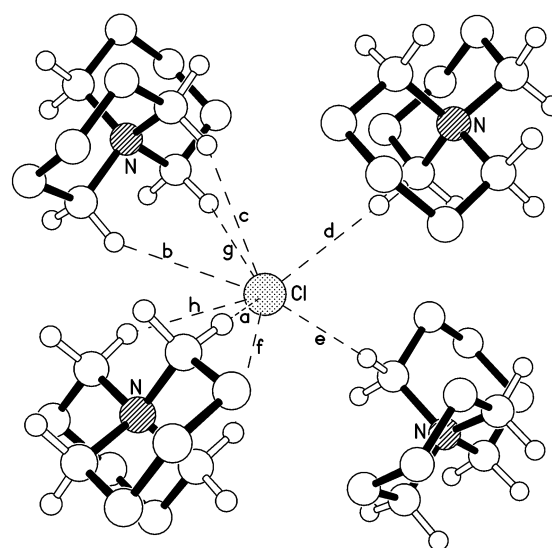


Fig. 2 X-ray crystal structure of [6+6]Cl, 5, demonstrating apparent C–H...Cl hydrogen-bonding interactions between chloride and neighbouring azoniaspiro cations (only α -H atoms are shown, for clarity).



suggested fundamental disorder in the crystals, which may indicate the presence of multiple cation structures arising from a high degree of conformational flexibility of the seven-membered rings.

Subsequently, B3LYP-D2/6-311++G(d,p) calculations were employed to locate low-energy ion pair conformers, initially for the single-ring salt $[C_4C_1\text{pip}]\text{Cl}$, **2**, and then for azoniaspiro salt $[7+6]\text{Cl}$, **8**. Cation structures were initially optimised, in the absence of the chloride anion, with differing combinations of ring conformation (chair and boat type). The chloride anion was then positioned in chemically-sensible locations around the lowest-energy cation structures, and ion pair optimisation calculations were performed. Ion pair structures and $\Delta E/\Delta G$ values are summarised below for **2** (Fig. 3, Table 1) and **8** (Fig. 4, Table 2).

A naming convention was chosen to describe the different ion pair conformers of **2**. Chair conformers are denoted 'C', and boat conformers as 'B'. Ion pairs of **2** were distinguished also according to whether the methyl/butyl chains were in the axial position ('Me-Ax' and 'Bu-Ax', respectively).

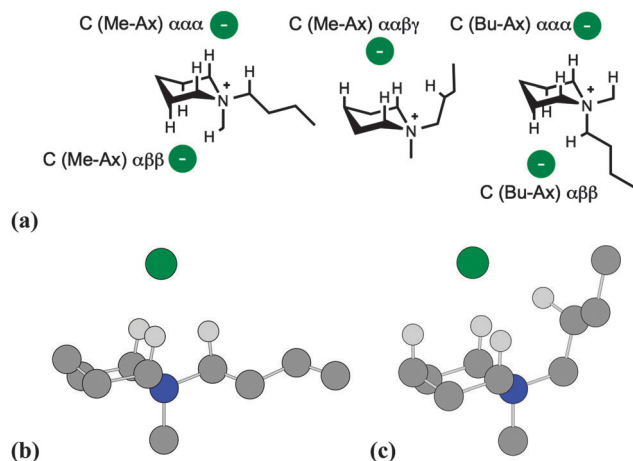


Fig. 3 (a) Primary cation–anion interactions for chair conformers of $[C_4C_1\text{pip}]\text{Cl}$, **2**. Structures of the: (b) 'C (Me-Ax) $\alpha\alpha\alpha$ 1', and; (c) 'C (Me-Ax) $\alpha\beta\beta\gamma$ ' ion pair conformers of **2**, showing primary H-bond interactions. All other hydrogen atoms are omitted, for clarity.

Table 1 Relative energies of chair ('C') and boat ('B') ion pair conformers of $[C_4C_1\text{pip}]\text{Cl}$, **2**. ΔE values are zero-point corrected

Ion pair $[C_4C_1\text{pip}]\text{Cl}$	ΔE (kJ mol ⁻¹)	ΔG (kJ mol ⁻¹)
C (Me-Ax) $\alpha\alpha\alpha$ 1	0.0	0.0
C (Bu-Ax) $\alpha\alpha\alpha$ 1	1.8	2.0
C (Bu-Ax) $\alpha\alpha\alpha$ 2	4.7	4.4
C (Me-Ax) $\alpha\alpha\alpha$ 2	5.1	5.2
C (Me-Ax) $\alpha\alpha\beta\gamma$	13.2	16.5
C (Bu-Ax) $\alpha\beta\beta$	17.4	19.2
C (Bu-Ax) $\alpha\alpha\beta$	21.6	21.1
B $\alpha\alpha\beta$ 1	23.9	24.3
B $\alpha\alpha\alpha$ 1	25.1	25.9
C (Me-Ax) $\alpha\beta\beta$	29.7	29.6
B $\alpha\alpha\alpha$ 2	30.3	31.8
B $\alpha\alpha\alpha$ 3	31.1	30.8
B $\alpha\alpha\alpha$ 4	31.3	31.2
B $\alpha\alpha\beta$ 2	46.3	45.5

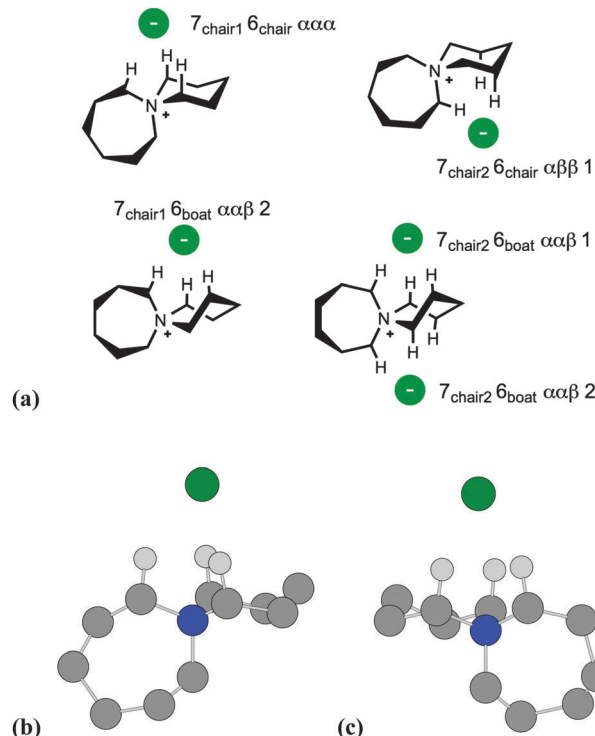


Fig. 4 (a) Diagram representing primary cation–anion interactions for $[7+6]\text{Cl}$, **8**. Structures of the low-energy (b) $7_{\text{chair}1} 6_{\text{chair}} \alpha\alpha\alpha$ 1' and (c) $7_{\text{chair}2} 6_{\text{chair}} \alpha\alpha\alpha$ 1' ion pair conformers of **8**, showing primary H-bond interactions. All other hydrogen atoms are omitted, for clarity.

Table 2 Relative energies of ion pair conformers of $[7+6]\text{Cl}$, **8**. ΔE values are zero-point corrected

Ion pair 6-azoniaspiro[6.5]-dodecanium chloride	ΔE (kJ mol ⁻¹)	ΔG (kJ mol ⁻¹)
$7_{\text{chair}1} 6_{\text{chair}} \alpha\alpha\alpha$	0.0	0.0
$7_{\text{chair}2} 6_{\text{chair}} \alpha\alpha\alpha$ 1	1.6	2.0
$7_{\text{chair}1} 6_{\text{chair}} \alpha\alpha\beta$ 1	4.3	4.0
$7_{\text{chair}2} 6_{\text{chair}} \alpha\alpha\alpha$ 2	4.7	4.1
$7_{\text{chair}1} 6_{\text{chair}} \alpha\alpha\beta$ 2	10.2	9.9
$7_{\text{chair}2} 6_{\text{boat}} \alpha\alpha\alpha$	24.7	23.8
$7_{\text{chair}2} 6_{\text{chair}} \alpha\beta\beta$ 1	25.9	23.9
$7_{\text{chair}2} 6_{\text{chair}} \alpha\beta\beta$ 2	26.0	24.2
$7_{\text{chair}2} 6_{\text{boat}} \alpha\alpha\beta$ 1	30.7	28.8
$7_{\text{chair}1} 6_{\text{boat}} \alpha\alpha\beta$ 1	35.9	33.9
$7_{\text{chair}1} 6_{\text{boat}} \alpha\alpha\beta$ 2	36.6	33.9
$7_{\text{chair}2} 6_{\text{boat}} \alpha\alpha\beta$ 2	38.2	35.7
$7_{\text{chair}2} 6_{\text{boat}} \alpha\beta\beta$	45.5	43.4

Furthermore, conformers were described in terms of the chloride anion location and the protons of the cation with which it interacts. Hence, the lowest-energy 'C (Me-Ax) $\alpha\alpha\alpha$ 1' ion pair conformer of **2** (Table 1) is a structure whereby the methyl group is in the axial position; the chloride anion interacts with three α -protons on the $[C_4C_1\text{pip}]^+$ cation. Eight ion pair geometries of **2** with a cation chair ('C') conformation were located, and six boat ('B') conformers were also optimised (Table 1). All boat conformers were higher in energy than chair equivalents, with the exception of the 'C (Me-Ax) $\alpha\beta\beta$ ' structure.



A similar naming convention was employed to distinguish different ion pair structures of [7+6]Cl, **8**. Two distinct chair-type conformers were observed for the seven-membered ring. The six-membered ring could adopt either a chair or boat conformation. In addition, the conformers were also described in terms of the aliphatic ring protons that interact closely with the chloride anion, as for compound **2**. Hence, for the lowest-energy '7_{chair1} 6_{chair} ααα' ion pair conformer of **8**, the seven-membered ring adopts the first of two chair conformers, and the six-membered ring also adopts a chair structure. Chloride then interacts with three α protons on the azoniaspiro cation.

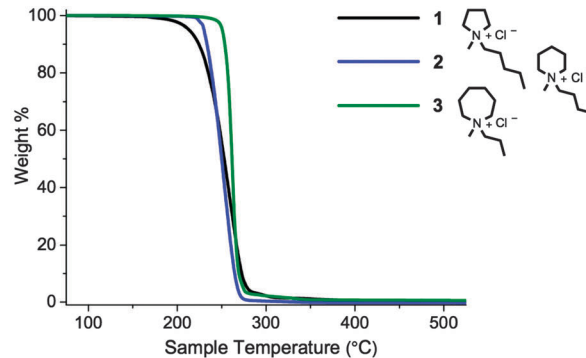
For the majority of ion pair conformers of **2**, the chloride anion interacts with a tripod array of α and β protons radiating out of the cation. Higher-energy conformers ($\Delta E = 13\text{--}30\text{ kJ mol}^{-1}$) involve bonding interactions of Cl[−] with β/γ protons on the piperidinium cation (C–H···Cl distances 2.36–2.49 Å).

The four lowest-energy ion pairs of salt **2** are very close in energy ($\Delta E = <6\text{ kJ mol}^{-1}$), and involve the interaction of the chloride anion with three α protons on the piperidinium cation. There appears to be no substantial energy increase associated with the 1,3-diaxial interactions of an axial butyl substituent.

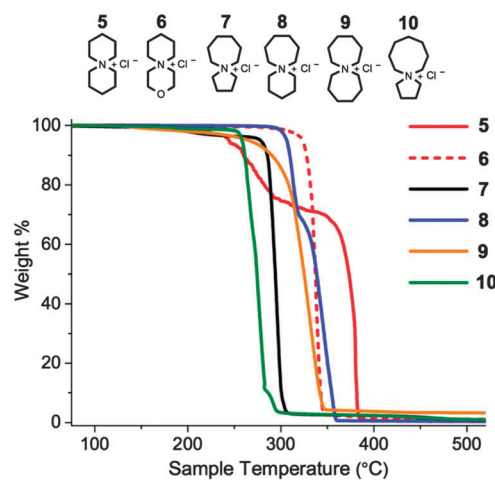
One particular ion pair structure of **2** ('C (Me-Ax) ααβγ') is approximately 13 kJ mol^{−1} higher in energy than the lowest-energy conformer and involves interaction of a butyl β proton (facilitated by –Bu rotation) and also a ring γ proton (Fig. 3c).

Ion pair conformers of salt **8** with 'boat'-conformation six-membered rings ($\Delta E = 24\text{--}46\text{ kJ mol}^{-1}$) typically lie higher in energy than those with 'chair'-conformation rings ($\Delta E = 0\text{--}26\text{ kJ mol}^{-1}$). For all optimised ion pairs of **8**, the chloride anion interacts preferentially with a tripodal arrangement of α and β protons in close proximity to the central quaternary nitrogen atom of the azoniaspiro cation (C–H···Cl distances 2.35–2.50 Å, cf. 2.75–2.95 Å for the X-ray structure of related [6+6]Cl, **5**, Fig. 2). The two lowest-energy ion pair conformers of salt **8**, '7_{chair1} 6_{chair} ααα' and '7_{chair2} 6_{chair} ααα', (Fig. 4(b) and (c)) are isoenergetic within the computational error, despite having different core cation geometries. Indeed, relative energies of the four lowest-energy conformers (Table 1) may become re-ordered upon a change in the computational method. The X-ray investigation of **8** demonstrated fundamental disorder for cation structure, possibly in accordance with the observed co-existence of analogous degenerate structures in the crystal-line material.

Therefore DFT calculations of [C₄C₁pip]Cl, **2** and [7+6]Cl, **8**, indicate that Cl[−] preferentially adopts a position close to the quaternary nitrogen atom, interacting with three protons in a tripodal array. Stable computed conformers of **8** closely match those observed in the X-ray structure of the analogous salt **5**. Furthermore the experimental (X-ray) C–H···Cl bond distances are not dissimilar from the calculated (DFT) distances. These observations therefore offer valuable insights into low-energy geometries of tetraalkylammonium salts, representing 'ground state' conformers when calculating thermal decomposition re-action profiles.



(a)



(b)

Fig. 5 Temperature-ramped TGA thermographs for: (a) single-ring (**1–3**), and; (b) azoniaspiro (**5–10**) chloride salts, measured with heating rate 10 °C min^{−1}, and shown in the range 75–525 °C for clarity.

Subsequently, thermal stabilities of salts **1–10** were compared by employing temperature-ramped Thermogravimetric Analysis (TGA). Thermographs are displayed in Fig. 5 and T_{onset} , T_{start} and melting point values (T_{m}) are listed in Table 3.

Spirocyclic tetraalkylammonium chloride salts (**5–10**) are, in general, more thermally stable than single-ring analogues (**1–3**), exhibiting higher T_{onset} and T_{start} values; the increased rigidity of the cation, and the increased steric crowding in the proximity of the quaternary nitrogen atom, may partly explain this observation. For single-ring salts **1–3**, a slight enhancement in

Table 3 T_{onset} , T_{start} and T_{m} temperatures for tetraalkylammonium chloride salts, **1–3**, **5–10**

Salt	T_{onset} (°C)	T_{start} (°C)	T_{m} (°C)
1 [C ₅ C ₁ pyrr]Cl	228	198	127–129
2 [C ₄ C ₁ pip]Cl	232	221	>210
3 [C ₃ C ₁ azp]Cl	254	241	244–245*
5 [6+6]Cl	245	189	310–312*
6 [3-oxa-6+6]Cl	330	310	>280*
7 [7+5]Cl	286	274	262–263*
8 [7+6]Cl	323	297	272–273*
9 [7+7]Cl	304	263	>305*
10 [8+5]Cl	263	253	256–257*

* = melting with decomposition.



thermal stability with increasing ring size was observed, from the five-membered pyrrolidinium salt **1** ($T_{\text{onset}} = 228\text{ }^{\circ}\text{C}$) to the seven-membered azepanium salt **3** ($T_{\text{onset}} = 254\text{ }^{\circ}\text{C}$). Within the investigated series of azoniaspiro species, **5**–**10**, a significant variation in thermal decomposition behaviour was observed. Salts incorporating five-membered rings (**7** and **10**) exhibited poorer thermal stabilities than the salts composed of six- and seven-membered rings.

Two of the investigated salts exhibit clear, multi-step weight loss curves: [6+6]Cl, **5**, and [7+6]Cl, **8**. For these two species, approximately 30 wt% is lost in the first step, followed by the remaining 70% in the second thermal event. This decomposition profile was also observed with a slower TGA scan rate of $2\text{ }^{\circ}\text{C min}^{-1}$; this indicates that the phenomenon is unlikely to be the result of a heat transfer effect. It is instead possible that two or more crystalline polymorphs exist in the solid material, accounting for the two distinct thermal degradation events.

For certain compounds in the series, the first appreciable loss of mass (T_{start} , $dw/dt \neq 0$) occurs at a similar temperature to the measured melting point (T_{m}). This is most apparent for [C₃C₁azp]Cl **3**, and for the two azoniaspiro salts incorporating a five-membered ring, **7** and **10**. The thermographs for these salts are also noticeably steeper than for other compounds in the series. This may indicate that decomposition of these salts is restricted by the reduced mobility of the ions until melting occurs; once the salt melts, decomposition occurs rapidly.

Two of the organic chloride salts, [7+6]Cl, **8**, and [3-oxa-6+6]Cl, **6**, display particularly high thermal stabilities. T_{onset} values for these salts (323 and 330 $^{\circ}\text{C}$, respectively) compare favourably with chloride ionic liquids in the chemical literature measured under similar experimental conditions; for example [C₄C₁im]Cl ($T_{\text{onset}} = 257\text{ }^{\circ}\text{C}$),³⁰ [C₁=C₂C₁im]Cl ($T_{\text{onset}} = 254\text{ }^{\circ}\text{C}$),³⁰ [(C₂)₄N]Cl ($T_{\text{onset}} = 264\text{ }^{\circ}\text{C}$).²⁰ Despite the prohibitively high melting points of azoniaspiro species **6** and **8**, and their lack of a stable-liquid range, these results nevertheless indicate that selected azoniaspiro cations confer a promising improvement in thermal stability. Exchanging chloride for alternative asymmetric, charge-diffuse anions (e.g. [OAc]⁻, [CH₃SO₄]⁻, [NTf₂]⁻) is likely to yield compounds with lower melting temperatures, yet that retain the beneficial property of high thermal stability imparted by the azoniaspiro cation.

One potential degradation product of **6** is the highly toxic oxygen mustard, bis(2-chloroethyl)ether, generated by attack of two chloride anions at the α -carbons of the morpholinium ring. The potential of generating this oxygen mustard presents a significant risk when handling this compound, and renders compound **6** an unattractive candidate for raised-temperature applications. Therefore on this basis, salt **6** was not investigated further. For this reason, azoniaspiro chloride salts based on thiomorpholine were also not studied. Instead, 6-azoniaspiro-[6.5]dodecanium chloride, [7+6]Cl, **8**, was identified to be the most thermally stable (and safe) azoniaspiro salt in the series, and was selected for further investigation.

It is well-established that the T_{onset} parameter, as obtained from temperature-ramped TGA evaluation, represents a profound overestimate of the long-term thermal stability of ionic

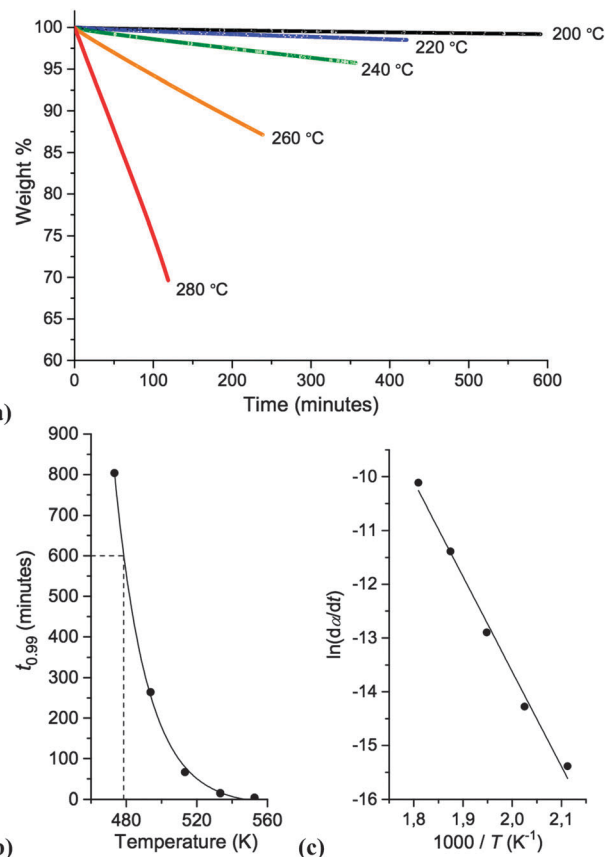


Fig. 6 Determination of the long-term thermal stability of [7+6]Cl, **8**: (a) isothermal TGA thermographs; (b) exponential fit of T (K) vs. $t_{0.99}$ (mins) to determine $T_{0.01/10}$; (c) determination of E_a and A .

liquids and related compounds; the ' $T_{0.01/10}$ ' value, proposed by MacFarlane and colleagues, represents the temperature at which exactly 1% decomposition occurs in 10 hours, and is a rudimentary estimate of the long-term upper operating temperature.^{21,22} $T_{0.01/10}$ values have been measured and compared for a broad spectrum of ionic liquids.^{21,22,28,30}

Isothermal TGA experiments were performed for [7+6]Cl, **8**. Graphs representing the long-term thermal stability of **8** are displayed in Fig. 6. Details of the isothermal TGA procedure for determining $T_{0.01/10}$ are described in the ESI.†

The $T_{0.01/10}$ value of [7+6]Cl, **8**, was measured as 205 $^{\circ}\text{C}$, substantially lower than both the T_{onset} (323 $^{\circ}\text{C}$) and T_{start} (297 $^{\circ}\text{C}$) measurements from temperature-ramped TGA in accordance with the chemical literature.²¹ As for T_{onset} the obtained $T_{0.01/10}$ value for **8** compares favourably with chloride ionic liquids in a previous comprehensive study by Cao and Mu; for example, [C₄C₁im]Cl ($T_{0.01/10} = 140\text{ }^{\circ}\text{C}$), [C₈C₁im]Cl ($T_{0.01/10} = 132\text{ }^{\circ}\text{C}$), [(HO)C₂C₁im]Cl ($T_{0.01/10} = 155\text{ }^{\circ}\text{C}$).³⁰ Hence, [7+6]Cl, **8**, can be employed for extended periods at temperatures over 50 $^{\circ}\text{C}$ higher than more traditional chloride-containing ionic liquids.

The observed linear isotherms for **8** indicate pseudo zeroth-order thermal decomposition, whereby one cation reacts with one anion to yield wholly volatile products, and thus the concentrations of ions do not change. In this circumstance,



Arrhenius parameters of thermal decomposition can be determined from the isothermal TGA data, according to eqn (1), below:³¹

$$d\alpha/dt = A \exp(-E_a/RT) \quad (1)$$

where $d\alpha/dt$ is the change in mole fraction, α , as a function of time. Plotting $\ln(d\alpha/dt)$ against $1/T$ yields a linear relationship with a gradient of $-E_a/R$ and a y-axis intercept $\ln(A)$ (Fig. 6c). E_a for thermal decomposition of [7+6]Cl, **8**, was measured as 147 kJ mol^{-1} , A as $2.7 \times 10^9 \text{ s}^{-1}$. These Arrhenius parameters enable rapid and accurate prediction of the rate of decomposition for **8** at different temperatures, when in an open system and under inert conditions.

Neither the temperature-ramped or isothermal TGA experiments yield any information pertaining to the precise chemical mechanisms that occur upon thermal decomposition of the chloride salts. However, TGA 'hyphenated' with Mass Spectrometry (TGA-MS) enables detection and characterization of volatile degradants; a flow of inert helium gas carries thermal decomposition products from the TGA apparatus, *via* a heated ceramic capillary, to be analysed by Electron Ionisation Mass Spectrometer (EIMS). At intervals, EI mass spectra are recorded, allowing the intensities of select m/z values to be plotted as a function of the sample temperature.

It is likely that EIMS will bring about the fragmentation of small-molecule thermal decomposition species. Therefore for certain products, the parent ion ($[M]^+$) may not necessarily be observed. The National Institute of Standards and Technology (NIST) Chemistry WebBook was employed in order to identify observed thermal decomposition products, based upon their expected fragmentation patterns.⁴⁶

Initially one single-ring compound, [C₄C₁pip]Cl **2**, and one azoniaspiro species, [7+6]Cl **8**, were investigated using TGA-MS. Experiments were performed on 20–60 mg of the salt and with the same heating rate of $10 \text{ }^\circ\text{C min}^{-1}$ employed for the regular temperature-ramped TGA study (Fig. 5). Helium was used as the carrier gas. Electron Ionisation (EI) mass spectra were collected at periodic intervals throughout the experiment in the range m/z 10–300 (ionisation energy 70 eV). The primary peaks of interest were of m/z 40 and above. Detailed TGA-MS procedures are described in the ESI.†

Plausible decomposition pathways were initially proposed for the single-ring salt [C₄C₁pip]Cl, **2** (Fig. 7). Chloride is a moderate nucleophile but a poor base; thermal decomposition mechanisms are therefore likely to involve action of chloride as a nucleophile (also in agreement with previous literature on halide ionic liquids³⁵). Decomposition may occur through S_N2 nucleophilic attack of the chloride anion at a carbon atom α to the quaternary nitrogen atom, bringing about a dealkylation of the methyl or butyl chains (Fig. 7, A1, B1 and C1) or opening of the six-membered ring (Fig. 7, D1 and E1). TGA-MS data for **2** is shown in graphical form (Fig. 8).

The TGA-MS analysis of **2** revealed a substantial peak, of m/z 50, evolving during the primary mass-loss event (*ca.* 225–300 °C); this peak was assigned to chloromethane, [CH₃Cl]⁺. In addition, smaller peaks of m/z 99 and m/z 141 were detected during the

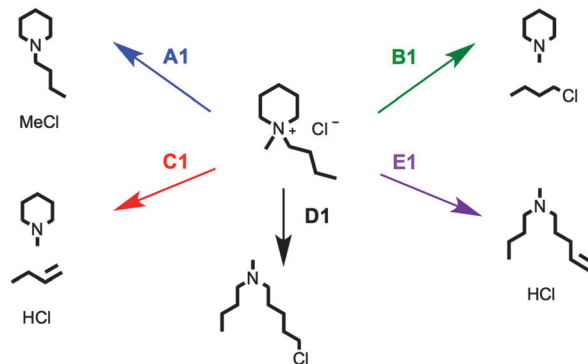


Fig. 7 Proposed S_N2 (A1, B1 and D1) and E2 Hofmann elimination (C1 and E1) thermal decomposition pathways for [C₄C₁pip]Cl, **2**.

main period of decomposition, corresponding to *N*-methylpiperidine [C₁pip]⁺ and *N*-butylpiperidine [C₄pip]⁺, respectively (intensities of peaks do not necessarily indicate relative quantities of degradation products, differences in ionisation efficiency/secondary ionisation should be considered). These three decomposition products strongly indicate that S_N2 nucleophilic substitution mechanisms occur at the [C₄C₁pip]⁺ methyl and butyl substituents (mechanisms A1 and B1, Fig. 7). By contrast, the TGA-MS investigation (Fig. 8) provided no clear evidence to suggest that E2 elimination mechanisms were simultaneously occurring during the thermal degradation of [C₄C₁pip]Cl, **2**.

The TGA-MS investigation of [7+6]Cl, **8**, did not provide any clear indication of specific thermal decomposition pathways; mass spectra obtained at regular intervals were far less trivial to interpret compared to salt **2**. No peak was observed at m/z 203, corresponding to the total mass of **8**, [C₁₁H₂₂NCl]⁺. A parent peak at this mass would suggest the opening of either the six or seven-membered ring by nucleophilic attack of the chloride anion, generating a neutral heterocyclic amine with a terminal chloroalkyl substituent. Furthermore, peaks for piperidine and azepane, which would form from complete detachment of one

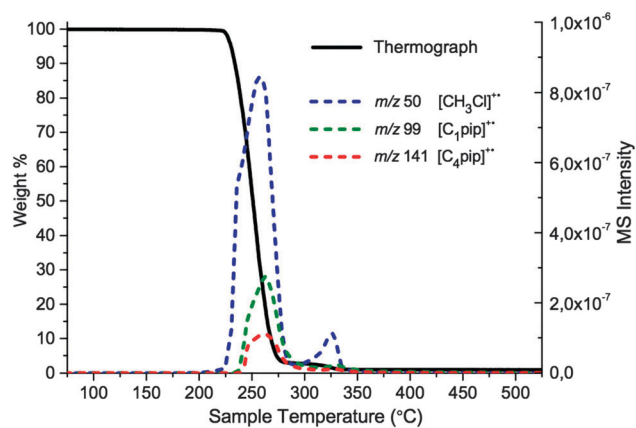


Fig. 8 TGA-MS evaluation of [C₄C₁pip]Cl, **2**. The continuous line denotes the TGA thermograph, the coloured dashed lines denote MS intensities for select m/z values. The assigned products are indicated. The thermograph is shown in the range 75–525 °C, for clarity.



of the ring alkyl substituents, were absent. However, the absence of the parent MS peak does not necessarily infer that such decomposition species do not form.

The added complexity of the thermal decomposition mechanisms of [7+6]Cl, **8**, demanded the use of Thermogravimetric Analysis-Gas Chromatography/Mass Spectrometry (TGA-GCMS). A primary advantage of the additional GC step is the separation of individual components in the product mixture. The difficulties associated with superimposed EI mass spectra (arising from the simultaneous evolution of all decomposition products) in TGA-MS experiments are therefore minimised.

TGA-GCMS experiments were performed on smaller quantities (*ca.* 2 mg) of [7+6]Cl, **8**, using porcelain crucibles, and a heating rate of 10 °C min⁻¹. Thermal decomposition products were transferred, *via* a heated capillary, and were collected at two temperatures during the ramp (310 and 340 °C) onto activated carbon tubes, for periods of 15 seconds. Subsequently, the tubes were loaded into the GCMS apparatus, and the tubes were heated to evolve the adsorbed decomposition products. Observed GC peaks ('i-viii', Fig. 9, Table 4) were cautiously assigned to thermal decomposition products by comparing the MS data against the NIST Chemistry WebBook.⁴⁶ Full TGA-GCMS procedures are described in the ESI,[†] including photos of the experimental setup (Fig. E4, ESI[†]). Proposed thermal degradation mechanisms of **8** are indicated in Fig. 10, based on the assigned products in Table 4.

The two chromatographs in Fig. 9, representing thermal decomposition products isolated at 310 and 340 °C, are highly similar; the majority of the significant peaks are present in both spectra. This observation indicates that the two separate mass loss steps in the temperature-ramped TGA thermograph of **8** do not correspond to two distinct chemical mechanisms. Instead, the two step weight loss phenomenon possibly arises due to a heat transfer effect associated with the melting of the salt, or due to multiple crystalline polymorphs of **8** that exhibit marginally differing thermal properties.

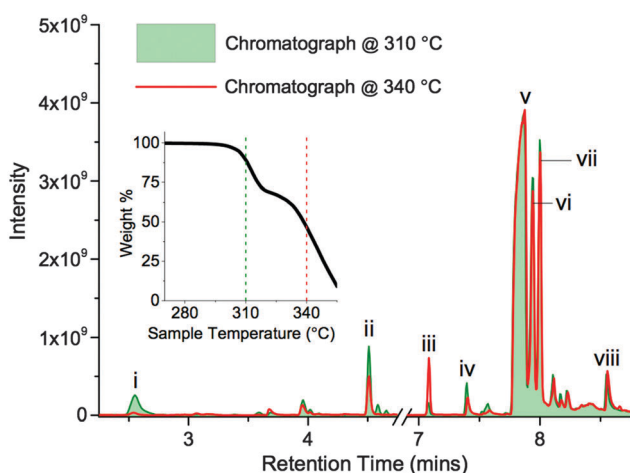


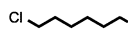
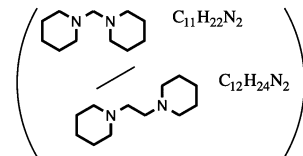


Fig. 9 Gas chromatograms of TGA-GCMS experiments, representing thermal decomposition of [7+6]Cl, **8**. Selected primary peaks are labelled (i–viii). The two overlaid chromatograms highlight products condensed onto activated carbon tubes, for 15 second periods, at *ca.* 310 and 340 °C in the TGA thermograph (inset).

Table 4 TGA-GCMS gas chromatograph peaks i–ix, for the thermal decomposition of [7+6]Cl, **8**. Products are tentatively assigned

Peak label	RT (min)	Assigned Product(s)
i	2.55	 C ₆ H ₁₀
ii	4.51	 C ₆ H ₁₁ Cl
iii	7.08	 C ₆ H ₁₂ Cl ₂
iv	7.41	 C ₁₁ H ₂₂ N ₂ C ₁₂ H ₂₄ N ₂
v	7.87	
vi	7.94	
vii	8.00	
viii	8.56	
ix	9.94	

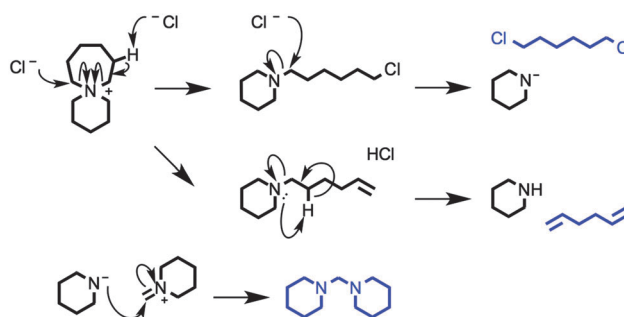


Fig. 10 A series of possible thermal decomposition mechanisms for 6-azoniaspiro[6.5]dodecenium chloride, [7+6]Cl, **8**. Degradation products, tentatively assigned in the TGA-GCMS investigation (Fig. 9, Table 4), are highlighted in blue.

One exception is peak 'i', at a retention time (RT) of 2.55 minutes, which is primarily absent from the chromatograph at 340 °C. The mass spectrum of this peak closely matches that of 1,5-hexadiene, [C₆H₁₀]⁺.⁴⁶ This degradation product may form by elimination reactions occurring sequentially at both β carbons of the seven-membered ring of **8** (Fig. 10). Another minor peak, at RT = 7.08 mins, closely matches (the MS data of) 1,6-dichlorohexane, [C₆H₁₂Cl₂]⁺, likely to be formed from sequential nucleophilic attack of two chloride anions at the two α carbons of the seven-membered ring of **8**.⁴⁶

The most prominent peaks in the TGA-GCMS study of **8** occur in the region RT = 7.8–8.1 mins. Mass spectra associated with these peaks each matched moderately well with two known compounds, bis(1-piperidyl)methane and 1,2-bis(1-piperidyl)ethane (Table 4). It is therefore possible that either or both of these piperidine-containing products are present in the thermal decomposition mixture, and correspond to particular peaks in the ensemble at RT *ca.* 8 mins. However, three major peaks (at RT = 7.87, 7.94 and 8.00 minutes) each display an intense fragment mass peak at *m/z* 98. This can be assigned to the 1-methylenepiperidinium cation radical [(CH₂)₅N=CH₂]⁺. This fragment at *m/z* 98 would likely be generated from any decomposition product containing the *N*-methylene-piperidine subunit, including the two open-chain intermediates generated



from S_N2 and/or E2 ring-opening mechanisms of the seven-membered ring; these intermediates, *N*-(6-chlorohexyl)piperidine and *N*-(hex-5-enyl)piperidine, are displayed in Fig. 10.

No reference EIMS data for these complex structures exists in the NIST Chemistry WebBook.⁴⁶ Therefore, it is possible that the significant gas chromatograph peaks (RT = 7.8–8.1 mins) correspond to these decomposition products from the opening of the seven-membered ring.

Although these primary peaks of the gas chromatographs cannot be assigned with certainty, the observations indicate that thermal decomposition mechanisms occur predominantly at the seven-membered ring of [7+6]Cl, **8**, leaving the six-membered piperidine ring intact. Increased conformational flexibility (and the unfavourable bond angles) of the seven-membered ring may partially account for this. Regardless of the explanation, azoniaspiro salts must undergo ring-opening upon thermal decomposition in direct contrast to the single-ring systems (*e.g.* salt **2**).

Subsequently, following the thermogravimetric investigation into thermal decomposition of single-ring and azoniaspiro chloride salts, B3LYP-D2/6-311++G(d,p) calculations were employed in order to rationalise observed degradation mechanisms of [C₄C₁pip]Cl, **2**, and [7+6]Cl, **8**. Details of the computational method are described in the ESI†

Initially, five transition state (TS) structures of [C₄C₁pip]Cl, **2**, corresponding to the five mechanisms in Fig. 7, were obtained. These TS structures are displayed below in Fig. 11. ΔE , ΔG , ΔH and $T\Delta S$ energies for the TSs and products are listed in Table 5, quoted in kJ mol⁻¹ relative to the lowest energy 'C (Me-Ax) $\alpha\alpha\alpha$ ' ion pair conformer (Fig. 3).

The S_N2 dealkylation (A1, B1) and S_N2 ring-opening (D1) mechanisms, for thermal decomposition of [C₄C₁pip]Cl **2**, are 26–36 kJ mol⁻¹ lower in activation energy ($\Delta H_{\text{activation}}$) than the two E2 elimination pathways (C1 and E1), consistent with the nucleophilic character of the chloride anion. By incorporating a more basic anion (*e.g.* [OAc]⁻) elimination mechanisms may become favoured over nucleophilic substitution. The calculations indicate only minor disparities (<5 kJ mol⁻¹) in $\Delta H_{\text{activation}}$ energy barriers for the multiple S_N2 mechanisms (A1, B1 and D1), and thus are of a similar magnitude to the computational error associated with this level of DFT calculation. Gas-phase calculations, therefore, find no difference in activation barrier between S_N2 pathways. The TGA-MS analysis suggested that mechanism A1 involving dealkylation at the *N*-methyl substituent might predominate (indicated by the more intense peak at *m/z* 50, [CH₃Cl]⁺, and the smaller peak at *m/z* 141, [C₄pip]⁺).

Calculations using the SMD solvation methodology⁴⁷ were subsequently carried out, enabling an approximation of the thermal degradation occurring within a solvent environment (IL). Solvent parameters related to the IL [C₄C₁pip]Cl were employed, and all structures were re-optimised and confirmed as minima or TS at the B3LYP-D2/6-311++G(d,p) level of theory (further details are provided in the ESI†). Similar to the gas-phase calculations, there are only minor differences in the $\Delta H_{\text{activation}}/\Delta G_{\text{activation}}$ barriers for the different S_N2 mechanisms

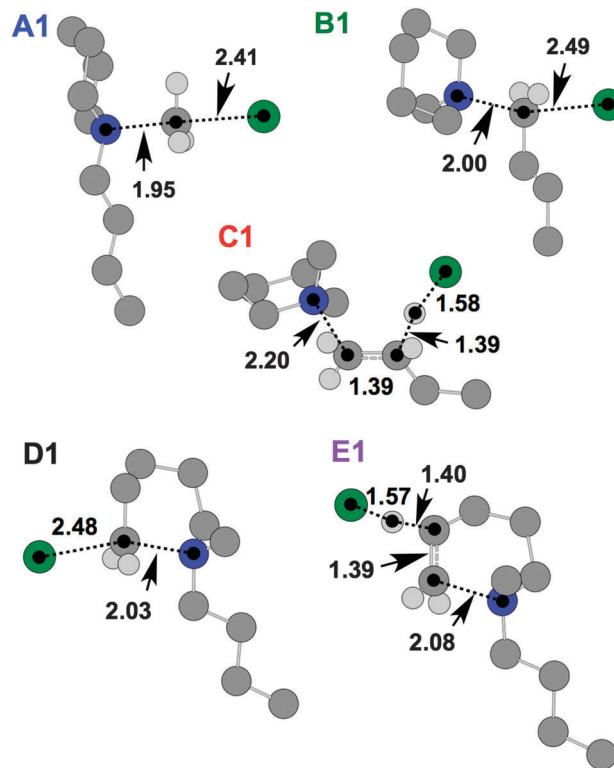


Fig. 11 TS structures for [C₄C₁pip]Cl, **2**, representing S_N2 (A1, B1 and D1) and E2 (C1 and E1) mechanisms. Atom distances are in angstroms.

Table 5 Gas phase TS and product energies for decomposition mechanisms of [C₄C₁pip]Cl, **2**, relative to the lowest energy 'C (Me-Ax) $\alpha\alpha\alpha$ ' ion pair conformer (Fig. 3). ΔE values are zero-point corrected

Mechanism	Transition state (kJ mol ⁻¹)				Product (kJ mol ⁻¹)			
	ΔE	ΔG	ΔH	$T\Delta S$	ΔE	ΔG	ΔH	$T\Delta S$
A1	88.4	86.7	89.0	2.3	13.4	-34.4	15.5	49.9
B1	90.2	91.3	90.0	-1.2	20.4	-31.2	21.2	52.4
C1	117.5	116.0	119.0	3.1	77.4	-12.4	83.0	95.4
D1	92.8	91.5	93.3	1.8	7.8	-3.5	12.1	15.5
E1	122.6	118.1	124.6	6.5	65.3	15.6	74.2	58.6

with SMD (Table E7, ESI†). Although uncertain (considering small TS $T\Delta S$ values, Table E7, ESI†), it is possible that a subtle entropic cost associated with Cl⁻ attack at the flexible butyl substituent is an explanation for the calculated difference in $\Delta G_{\text{activation}}$ for mechanisms A1 and B1.

E2 mechanisms (C1 and E1) involve large positive entropy changes ($T\Delta S = +95.4$ and $+58.6$ kJ mol⁻¹ respectively (gas phase), $+89.7$ and $+53.0$ kJ mol⁻¹ respectively (IL), Table 5), associated with formation of two or three neutral molecules from one ion pair.

TS and product geometries were also investigated for the thermal decomposition mechanisms of [7+6]Cl, **8**, using DFT (Fig. 12, Table 6). The TGA-GCMS investigation of **8** (Fig. 9 and 10) lent evidence towards preferential opening of the seven-membered ring. In an attempt to rationalise this observation, $\Delta G_{\text{activation}}$ barriers for S_N2 and E2 ring-opening mechanisms of both the six and seven-membered rings were compared.



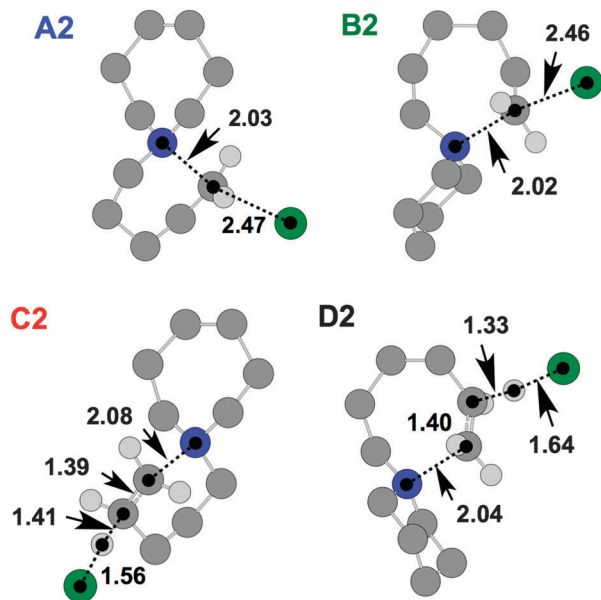


Fig. 12 TS structures for [7+6]Cl, **8**, representing S_N2 (A2 and B2) and E2 (C2 and D2) mechanisms relative to the lowest energy ${}^7\text{chair}_1\text{6chair}\ \alpha\alpha\alpha'$ ion pair conformer (Fig. 4). Atom distances are in angstroms.

Table 6 Gas phase TS and product energies for decomposition mechanisms of [7+6]Cl, **8**, relative to the lowest energy ${}^7\text{chair}_1\text{6chair}\ \alpha\alpha\alpha'$ ion pair conformer. ΔE values are zero-point corrected

Mechanism	Transition state (kJ mol^{-1})				Product (kJ mol^{-1})			
	ΔE	ΔG	ΔH	$T\Delta S$	ΔE	ΔG	ΔH	$T\Delta S$
A2	97.4	96.0	97.9	1.9	9.6	-2.4	14.0	16.4
B2	93.3	93.3	93.3	0.0	-8.4	-21.8	-4.4	17.4
C2	127.6	122.4	129.5	7.1	67.1	16.5	76.2	59.6
D2	115.3	109.8	117.1	7.3	50.0	-2.9	59.3	62.2

The calculations of thermal decomposition mechanisms of $[\text{C}_4\text{C}_1\text{pip}]\text{Cl}$, **2**, demonstrated that $\Delta G_{\text{activation}}$ barriers for S_N2 mechanisms were 24–32 kJ mol^{-1} (gas phase) or 37–55 kJ mol^{-1} (IL) lower in energy than for the E2 mechanisms. Similarly, for [7+6]Cl, **8**, $\Delta G_{\text{activation}}$ energies for S_N2 pathways (A2, B2) are approximately 14–29 kJ mol^{-1} (gas phase) or 6–38 kJ mol^{-1} (IL) lower than for the corresponding E2 mechanisms (C2, D2).

Moreover, routes involving opening of the seven-membered ring are slightly more facile ($\Delta\Delta G_{\text{activation}} = 3\text{--}13\ \text{kJ mol}^{-1}$ (gas phase) or 4–28 kJ mol^{-1} (IL)) than those that involve the opening of the six-membered ring, consistent with the findings of the TGA-GCMS experiments (Fig. 9 and 10). These minor disparities in $\Delta G_{\text{activation}}$ suggest that E2 mechanisms could predominate if the chloride anion were to be substituted for other anions. Indeed, the precise thermal decomposition mechanisms that predominate for all single-ring/azoniaspiro tetraalkylammonium salts are likely to be heavily dependant on the nature of the anion, as for the dialkylimidazolium ionic liquids²⁰ and others³⁰ in the chemical literature.

The calculated activation barrier for each of the investigated mechanisms of **8** ($\Delta H_{\text{activation}} = 93\text{--}130\ \text{kJ mol}^{-1}$ (gas phase)) are substantially lower than the experimental activation energy ($E_a = 147\ \text{kJ mol}^{-1}$) from isothermal TGA experiments.

An explanation for this disparity may be due, partially, to the limitations of using single ion pairs (stable clusters of one cation and one anion) as the starting point for thermal decomposition reaction profiles. Recent investigations have demonstrated that ion pair dimers are more capable of recovering the structural and bonding features that are exhibited in the bulk ionic liquid (or crystalline phase).^{43–45} The X-ray structure of azoniaspiro compound [6+6]Cl, **5**, Fig. 2 (analogous to [7+6]Cl, **8**) reveals that any particular chloride anion is likely to be strongly interacting with α and β hydrogen atoms on as many as four neighbouring cations. Therefore, an additional input of energy will be required to break these favourable interactions between chloride and the near cations, moreover chloride will be less nucleophilic in this more ‘solvated’ environment.

Calculations using the SMD model⁴⁷ were carried out for [7+6]Cl, **8** (Table E8, ESI[†]), similarly to the calculations performed for **2**. Incorporation of the SMD solvation model appears to recover a large proportion of the energy disparity between the calculated gas-phase $\Delta H_{\text{activation}}$ values and the experimental E_a from isothermal TGA (147 kJ mol^{-1}). With the addition of SMD, $\Delta H_{\text{activation}}$ was calculated as +140.0 and +135.7 kJ mol^{-1} for S_N2 mechanisms A2 and B2 respectively. Similar to gas-phase calculations, solvated E2 pathways exhibited $\Delta H_{\text{activation}}$ values significantly higher ($\Delta\Delta H_{\text{activation}} = 13\text{--}45\ \text{kJ mol}^{-1}$) than for the corresponding S_N2 pathways. The SMD results therefore indicate that it is critical to account for solution effects when calculating transition state energy barriers for thermal decomposition routes.

Comparing T_{onset} values for azoniaspiro chloride salts **4–10** (Table 3), it is sensible to suggest that six and seven-membered rings impart improved thermal stability over five or eight-membered rings. Furthermore, TGA-GCMS and DFT investigations into **8** reveal that decomposition occurs more readily at the seven-membered ring. On this basis, [6+6]Cl, **5**, should exhibit the greatest thermal stability of the compounds in the series. However, **5** was found to be among the most unstable, degrading *via* two distinct thermal events ($T_{\text{start}} = 189\ ^\circ\text{C}$). To minimise the possibility that an impurity is responsible for the poor observed stability of **5**, repeat temperature-ramped TGA experiments were performed, and a separate batch of the salt was prepared (omitting MgSO_4 in the purification); multi-step thermographs were obtained in each circumstance. Moreover, spectroscopic analysis of **5** confirmed high purity. Therefore, the low thermal stability of [6+6]Cl, **5**, cannot reasonably be attributed to the presence of an impurity.

Nevertheless the focal azoniaspiro salt in the investigation, [7+6]Cl, **8**, exhibited significantly improved thermal stability relative to single-ring analogues **1–3**; in terms of the T_{onset} and T_{start} parameters, **8** outperformed $[\text{C}_4\text{C}_1\text{pip}]\text{Cl}$, **2**, by 91 and 76 $^\circ\text{C}$, respectively. The prohibitively high melting temperatures of the azoniaspiro chlorides presents a problem for employing these salts in ionic liquid applications; no salt in the series (**5–10**) exhibits a significant stable-liquid range. Nevertheless, substituting chloride for a larger, charge-diffuse molecular anion will lower the melting point substantially, whilst retaining the inherent stability imparted by the cation. To this end,



two additional azoniaspiro salts were prepared, by the reaction of **8** with either Li[OTf] or Li[Ntf₂] ([OTf][−] = trifluoromethanesulfonate, [Ntf₂][−] = bis(trifluoromethanesulfonyl)imide). The resultant salts, [7+6][OTf] and [7+6][Ntf₂], exhibited melting temperatures of 115.5–117 °C and 104.5–106 °C, respectively (full synthesis/characterisation is listed in the ESI†). The *T*_{onset} value of [7+6][OTf] was measured as 412 °C, comparing favourably or similarly with analogous heterocyclic triflate salts (e.g. [C₄py][OTf] *T*_{onset} = 362 °C;³⁰ [(HO)²C₂C₁im][OTf] *T*_{onset} = 391 °C;³⁰ [C₁C₁C₁⁴pz][OTf] *T*_{onset} = 395 °C⁴⁸ or; [C₄C₁im][OTf] *T*_{onset} = 393 °C³⁰ or 410 °C⁴⁸). The *T*_{onset} value of [7+6][Ntf₂] was measured as 431 °C, approximately equal to other ionic compounds incorporating this anion^{28,30} (degradation is known to proceed *via* an anion-centred degradation mechanism²⁷). Nevertheless, in contrast to the commonly-encountered dialkyl-imidazolium species, azoniaspiro salts are incapable of undergoing N-Heterocyclic Carbene (NHC) formation/metal chelation in the presence of reactive (dissolved) species. This may render them a more suitable choice of solvent media for organic synthesis and catalysis.

Preparation of binary/reciprocal binary mixtures^{48–52} of azoniaspiro species is likely to afford still lower melting points. Processes that exploit the specific interactions of a particular anion, therefore, may be undertaken, whilst problems associated with high reactivity between the cation and anion (leading towards decomposition) are minimised. Azoniaspiro species therefore represent worthy candidates for future investigation.

Conclusions

In summary, a broad ensemble of single-ring and azoniaspiro-type chloride organic salts were prepared, in both high purity and acceptable yields. Investigations were undertaken in order to gain insights into the structural and bonding properties, and to ascertain information on the trends, thermodynamics and underlying mechanisms of thermal decomposition.

X-ray crystallography and DFT studies demonstrated that, regardless of whether the cation incorporates a single ring, or two cyclic motifs, the chloride anion is preferentially located in close proximity to the cation quaternary nitrogen atom, interacting strongly with tripod-like ensembles of α and β protons. Interestingly, for the monohydrate of **4**, [5+5]Cl·H₂O, the water molecule appears to compete with and displace the chloride anion, occupying the position closest to the nitrogen.

The relative thermal stabilities of the chloride species were compared by employing temperature-ramped TGA and trends were elucidated. Single-ring salts were susceptible to thermal degradation at markedly lower temperatures than their azoniaspiro analogues. Possible explanations include higher rigidity, and the absence of flexible alkyl chains for azoniaspiro salts; this is supported by the observation that thermal decomposition of 1-butyl-1-methylpiperidinium chloride, [C₄C₁pip]Cl, **2**, principally involves S_N2 dealkylation at the pendant *N*-methyl and *N*-butyl chains. The absence of these alkyl chains for azoniaspiro salts implies that thermal degradation must proceed

via a less optimal ring-opening transition state; TGA-GCMS analysis of [7+6]Cl, **8**, indicated that primary decomposition products retain an intact six-membered piperidine ring. To a minor extent full detachment of chloroalkyl/alkenyl fragments (e.g. 1-chlorohex-5-ene, 1,5-hexadiene) occurs. DFT calculations supported these experimental findings, highlighting that mechanisms to open the seven-membered ring of **8** proceed through a lower-energy transition state than for opening of the six-membered ring.

Within the sequence of azoniaspiro chloride salts, thermal stability trends may be established, on the basis of the sizes of the incorporated rings. Six- and seven-membered rings impart greater thermal stability than five- and eight-membered rings. Foremost, compounds based on five-membered rings should be avoided on account of the poor thermal stabilities and their observed hygroscopic nature. Salts derived from morpholine and thiomorpholine should similarly be avoided, because of their potential to evolve toxic mustard species upon thermal decomposition. Future research is necessary, to determine the impact of ring-substitution and branching on thermal stability.

A relative lack of conformational flexibility of the azoniaspiro-cation salts has an unwanted side effect; melting temperatures are significantly raised. The reactive chloride anion was selected, in this investigation, for the deliberate purposes of incurring thermal decomposition (to compare relative stabilities of the different cation structures), to minimise the computational complexity/expense, and for the ease of laboratory preparation. For the examined azoniaspiro chloride salts, the onset of melting appears closely associated with thermal decomposition. However, pairing an azoniaspiro cation with a large, charge-diffuse and bulky anion (e.g. [OTf][−] or [Ntf₂][−]) yields compounds that have much lower melting points, albeit considerably above room temperature.

We propose that it is more sensible to consider the concept of a ‘stable-liquid range’; a temperature-range in which the species is both in the liquid state and is thermally stable for a prolonged period of time. ‘Room-Temperature Ionic Liquids’ (RTILs), though (by definition) molten at ambient conditions, often suffer from poor thermal stability.^{36,37} By contrast all-inorganic ‘Molten Salts’ are highly thermally-robust, yet their prohibitively high melting points preclude their use in a broad range of applications. Simple azoniaspiro salts, or more likely binary/reciprocal binary azoniaspiro mixtures,^{48,49} have the potential to partially bridge this ‘gap’ in stable-liquid ranges; difficulties associated with the high melting temperatures are therefore minimised, whilst retaining benefits associated with the high thermal stability of the azoniaspiro cation. Certain azoniaspiro chloride salts have been shown, in this investigation, to offer substantially higher operational temperature limits than single-ring (e.g. dialkylpiperidinium) equivalents.

Further investigation is required in order to gain a comprehensive understanding into thermal properties of azoniaspiro salts. Nevertheless, they represent promising candidate solvent systems, offering the beneficial properties of high-thermal stability, wide electrochemical window,¹⁸ and trivial synthesis and purification. Furthermore, a broad range of structures and



properties are achievable *via* the modification of the size and substituents of the two cyclic moieties, in analogy to the reputation of ionic liquids as ‘designer solvents’.

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References

- J. P. Hallett and T. Welton, *Chem. Rev.*, 2011, **111**, 3508–3576.
- F. H. Hurley and T. P. Wier, *J. Electrochem. Soc.*, 1951, **98**, 207–212.
- H. L. Chum, V. R. Koch, L. L. Miller and R. A. Osteryoung, *J. Am. Chem. Soc.*, 1975, **97**, 3264–3265.
- J. S. Wilkes, J. A. Levisky, R. A. Wilson and C. L. Hussey, *Inorg. Chem.*, 1982, **21**, 1263–1264.
- A. a. K. Abdul-Sada, A. M. Greenway, K. R. Seddon and T. Welton, *Org. Mass Spectrom.*, 1989, **24**, 917–918.
- A. a. K. Abdul-Sada, A. M. Greenway, K. R. Seddon and T. Welton, *Org. Mass Spectrom.*, 1992, **27**, 648–649.
- A. a. K. Abdul-Sada, A. M. Greenway, K. R. Seddon and T. Welton, *Org. Mass Spectrom.*, 1993, **28**, 759–765.
- J. S. Wilkes and M. J. Zaworotko, *J. Chem. Soc., Chem. Commun.*, 1992, 965–967, DOI: 10.1039/c39920000965.
- I. Anugwom, P. Maki-Arvela, P. Virtanen, P. Damlin, R. Sjoholm and J.-P. Mikkola, *RSC Adv.*, 2011, **1**, 452–457.
- D. R. MacFarlane, P. Meakin, J. Sun, N. Amini and M. Forsyth, *J. Phys. Chem. B*, 1999, **103**, 4164–4170.
- H. Sakaebe and H. Matsumoto, *Electrochem. Commun.*, 2003, **5**, 594–598.
- T. Belhocine, S. A. Forsyth, H. Q. N. Gunaratne, M. Nieuwenhuyzen, P. Nockemann, A. V. Puga, K. R. Seddon, G. Srinivasan and K. Whiston, *Green Chem.*, 2011, **13**, 3137–3155.
- F. F. Blicke and E. B. Hotelling, *J. Am. Chem. Soc.*, 1954, 5099–5103.
- A. Losurdo, W. Y. Wen and C. Jolicoeur, *J. Solution Chem.*, 1979, **8**, 449–460.
- R. Millini, L. Carluccio, F. Frigerio, W. O. Parker and G. Bellussi, *Microporous Mesoporous Mater.*, 1998, **24**, 199–211.
- F. Frigerio, L. Carluccio, W. O. Parker and R. Millini, *J. Mol. Catal. A: Chem.*, 2001, **166**, 167–174.
- S. I. Zones, A. W. Burton, G. S. Lee and M. M. Olmstead, *J. Am. Chem. Soc.*, 2007, **129**, 9066–9079.
- Y. Abu-Lebdeh, E. Austin and I. J. Davidson, *Chem. Lett.*, 2009, **38**, 782–783.
- M. R. R. Prasad and V. N. Krishnamurthy, *Thermochim. Acta*, 1991, **185**, 1–10.
- H. L. Ngo, K. LeCompte, L. Hargens and A. B. McEwen, *Thermochim. Acta*, 2000, 357–358, 97–102.
- C. P. Fredlake, J. M. Crosthwaite, D. G. Hert, S. Aki and J. F. Brennecke, *J. Chem. Eng. Data*, 2004, **49**, 954–964.
- T. J. Wooster, K. M. Johanson, K. J. Fraser, D. R. MacFarlane and J. L. Scott, *Green Chem.*, 2006, **8**, 691–696.
- A. Seeberger, A.-K. Andresen and A. Jess, *Phys. Chem. Chem. Phys.*, 2009, **11**, 9375–9381.
- F. Heym, B. J. M. Etzold, C. Kern and A. Jess, *Phys. Chem. Chem. Phys.*, 2010, **12**, 12089–12100.
- N. Meine, F. Benedito and R. Rinaldi, *Green Chem.*, 2010, **12**, 1711–1714.
- A. W. T. King, A. Parviainen, P. Karhunen, J. Matikainen, L. K. J. Hauru, H. Sixta and I. Kilpelainen, *RSC Adv.*, 2012, **2**, 8020–8026.
- Y. Chen, Y. Cao, Y. Shi, Z. Xue and T. Mu, *Ind. Eng. Chem. Res.*, 2012, **51**, 7418–7427.
- M. T. Clough, K. Geyer, P. A. Hunt, J. Mertes and T. Welton, *Phys. Chem. Chem. Phys.*, 2013, **47**, 20480–20495.
- N. E. A. Cousens, L. J. Taylor Kearney, M. T. Clough, K. R. J. Lovelock, R. G. Palgrave and S. Perkin, *Dalton Trans.*, 2014, **43**, 10910–10919.
- Y. Cao and T. Mu, *Ind. Eng. Chem. Res.*, 2014, **53**, 8651–8664.
- J. H. Flynn and L. A. Wall, *J. Polym. Sci., Part B: Polym. Lett.*, 1966, **4**, 323–328.
- B. Chan, N. Chang and M. Grimmitt, *Aust. J. Chem.*, 1977, **30**, 2005–2013.
- M. C. Kroon, W. Buijs, C. J. Peters and G.-J. Witkamp, *Thermochim. Acta*, 2007, **465**, 40–47.
- H. Ohtani, S. Ishimura and M. Kumai, *Anal. Sci.*, 2008, **24**, 1335–1340.
- S. D. Chambreau, J. A. Boatz, G. L. Vaghjiani, C. Koh, O. Kostko, A. Golan and S. R. Leone, *J. Phys. Chem. A*, 2012, **116**, 5867–5876.
- C. Maton, N. De Vos and C. V. Stevens, *Chem. Soc. Rev.*, 2013, **42**, 5963–5977.
- M. T. Clough, *Ionic Liquid Thermal Stability: implications for cellulose regeneration*, PhD thesis, Imperial College London, 2015.
- W. Eschweiler, *Ber. Dtsch. Chem. Ges.*, 1905, **38**, 880–882.
- P. A. Hunt and I. R. Gould, *J. Phys. Chem. A*, 2006, **110**, 2269–2282.
- P. A. Hunt, B. Kirchner and T. Welton, *Chem. – Eur. J.*, 2006, **12**, 6762–6775.
- P. A. Hunt, *J. Phys. Chem. B*, 2007, **111**, 4844–4853.
- P. A. Hunt, I. R. Gould and B. Kirchner, *Aust. J. Chem.*, 2007, **60**, 9–14.
- R. P. Matthews, C. Ashworth, T. Welton and P. A. Hunt, *J. Phys.: Condens. Matter*, 2014, **26**, 284112.
- R. P. Matthews, T. Welton and P. A. Hunt, *Phys. Chem. Chem. Phys.*, 2014, **16**, 3238–3253.
- R. P. Matthews, T. Welton and P. A. Hunt, *Phys. Chem. Chem. Phys.*, 2015, **17**, 14437–14453.



- 46 National Institute of Standards and Technology (NIST) Chemistry WebBook, <http://webbook.nist.gov/chemistry>.
- 47 A. V. Marenich, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. B*, 2009, **113**, 6378–6396.
- 48 M. T. Clough, C. R. Crick, J. Grasvik, P. A. Hunt, H. Niedermeyer, T. Welton and O. P. Whitaker, *Chem. Sci.*, 2015, **6**, 1101–1114.
- 49 H. Niedermeyer, J. P. Hallett, I. J. Villar-Garcia, P. A. Hunt and T. Welton, *Chem. Soc. Rev.*, 2012, **41**, 7780–7802.
- 50 J. N. Canongia Lopes, T. C. Cordeiro, J. M. S. S. Esperança, H. J. R. Guedes, S. Huq, L. P. N. Rebelo and K. R. Seddon, *J. Phys. Chem. B*, 2005, **109**, 3519–3525.
- 51 M. Y. Lui, L. Crowhurst, J. P. Hallett, P. A. Hunt, H. Niedermeyer and T. Welton, *Chem. Sci.*, 2011, **2**, 1491–1496.
- 52 M. Brussel, M. Brehm, A. S. Pensado, F. Malberg, M. Ramzan, A. Stark and B. Kirchner, *Phys. Chem. Chem. Phys.*, 2012, **14**, 13204–13215.

